

$m(\text{Cisob.}) = 3,52$

$n(\text{CO}_2) = \frac{5,772}{22,4} = 0,25 \text{ mol}$

$n(\text{C}_3\text{H}_8) = \frac{n(\text{CO}_2)}{11} = 0,022 \text{ mol}$

$n(\text{C}_4\text{H}_{10}) = \frac{n(\text{CO}_2)}{5,5} = 0,045 \text{ mol}$

$m(\text{C}_3\text{H}_8) = n \cdot M = 0,022 \text{ mol} \cdot 44 \text{ g/mol} = 0,92$

$m(\text{C}_4\text{H}_{10}) = n \cdot M = 0,045 \text{ mol} \cdot 58 \text{ g/mol} = 2,62 \text{ (gas)}$

$n(\text{C}_3\text{H}_8) = 0,022 - 33\%$

$n(\text{C}_4\text{H}_{10}) = 0,045 - 67\%$

Cisob. $0,045 - 67\%$

45

3,5

① Элемент X: P (фосфор) — элемент чужда.

② A — P_{сер.} 0,3

Б — P₂O₅ 0,3

В — P_{крас.} 0,3

Г — PH₄↑

Д — Na₃P₂

Е — HPO₃ 0,3

Ж — H₃PO₄ 0,3

З — Na₃PO₄

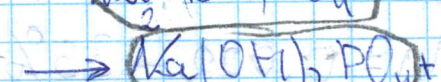
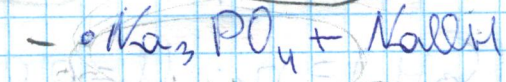
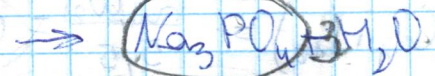
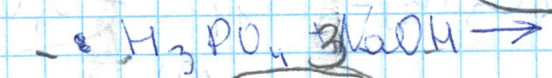
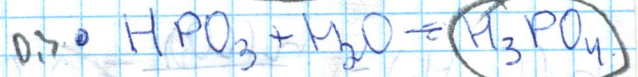
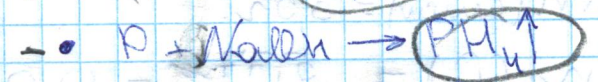
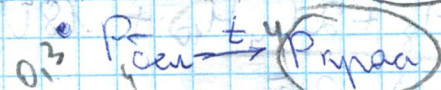
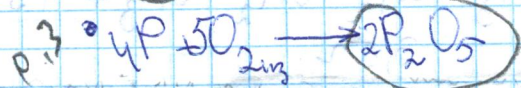
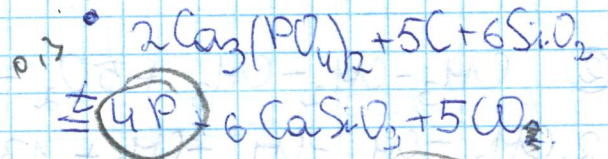
И — Na₂(OH)PO₄ 0,3

К — Na(OH)₂PO₄ 0,3

Л — PCl₃ 0,3

М HPCl₂ 0,3

③ Теория



Задание 10.3

а) 4, 7, 10, 1, 16, 15
 б) 2, 6, 9, 10, 0, 1, 50

$$\textcircled{8} \begin{array}{cccc|cc} 4 & 7 & 3 & 1 & 185 & 10 \\ + & + & + & + & & \end{array}$$

$$\textcircled{2} \begin{array}{cccc|cc} 2 & 9 & 1 & 10 & 25 & \textcircled{85} \\ + & + & + & + & & \end{array}$$

$$\textcircled{9} \begin{array}{cccc|cc} 1 & 7 & 8 & 10 & 0.5 & 25 \\ + & + & + & + & & \end{array}$$

3. Jaga 10.4

①

$$\frac{340 \text{ g NO}_3 \text{ (pp)}}{20\%} = \frac{340 \text{ g SO}_4 \text{ (p-p)}}{15\%}$$

$$\textcircled{2} m(\text{kg NO}_3) = \frac{20 \cdot 340}{100} = 68 \text{ t (20\%)}$$

$$m(\text{H}_2\text{O}) = 340 - 68 = 272 \text{ t (bagal)}$$



$$w(\text{kg NO}_3) = \underline{10,49\%}$$

③ Cmas: (masa reaksi)

$$10,49 (\text{kg NO}_3) - \textcircled{31,872} \text{ (masa reaksi awal)}$$

$$89,51 (\text{H}_2\text{O}) - 272$$

$$\underline{m(\text{p-p}) = 303,82}$$

④

$$\frac{303,82 \text{ g NO}_3 \text{ p-p}}{\quad} = \frac{303,82 \text{ g gasas masas}}{\text{Hg SO}_4 \text{ p-p}}$$

3,05



$m(\text{HNO}_3) = m(\text{MgSO}_4) = 303,87$

Substansi reagen:

$m(MgSO_4) = 51,82$ 0,5

$m(H_2O) = 289,2$

Stoek reagen

$m(MgSO_4) = 303,87 - 289,2 = 14,82$ — Osmalahko
 Inomoro $MgSO_4$

$51,8 - 15\%$

$14,8 - 4,2\%$

Omben $m(MgSO_4) = 4,2\%$

$51,8 - 14,8 = 37,2$ — yuno

$29,8 + 37 = 66,82$

Omben $marka = 66,82$

15

	HNO ₃	NaCl	Zagornu 10 5 Na ₃ PO ₄	HgNO ₃
1	✓	X	X	X
2	X	X	X	✓
3	X	✓	X	
4	X	X	✓	X

10.8.

1) продукта - HNO_3 2) продукта - AgNO_3 3) продукта Kall 4) продукта Na_3PO_4 